

# Chapter 6 Chemical Bonding Test

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CHAPTER 6 TEST: CHEMICAL BONDING REVIEW SHEET. 1. What are the 2 main types of bonds that we have learned about? ... What levels of differences in electronegativity correspond to polar covalent, nonpolar covalent, and ionic bonds? 6. If Sodium, Na, was to bond with Chlorine, Cl, what type of bond would they form? ...

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Chapter 6: Chemical Bonding. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. meguarino24. Terms in this set (62) Chemical bond. Mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together. Potential energy.

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CHAPTER 6 REVIEW Chemical Bonding SECTION 5 SHORT ANSWER Answer the following questions in the space provided. 1. Identify the major assumption of the VSEPR theory, which is used to predict the shape of atoms. Pairs of valence electrons repel one another. 2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn't water a linear molecule?

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Chemical Bonding Test: Subject: Chemistry General: Test Type: Mcqs: Total Question: 15: Total Marks: 30: Total Time: 20 Minutes: Given on this page is a free online quiz which includes important questions and answers about the Chemical Bonding. All the individuals who want to improve their general knowledge or are currently preparing for any ...

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preceding the word or expression that best completes the statement. 1. A chemical bond results from the mutual attraction of the nuclei of atoms for (a) electrons;

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CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

### **Chapter 6 Chemical Bonding Test Answer Key**

Chapter 6:Chemical Bonding-Test Review 1. List all the symbols of the elements on the periodic table that have a stable electron configuration (aka- a full outer shell). 2. In an electron dot diagram, what does the element symbol represent?

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A chemical bond that forms when atoms share electrons is always a(n) a. polar bond. c. metallic bond. b. ionic bond. d. covalent bond. \_\_\_\_ 5. When two fluorine atoms share a pair of electrons, the bond that forms is a(n) a. polar covalent bond. c. nonpolar covalent bond. b. ionic bond. d. double bond. \_\_\_\_ 6. The chemical formula for magnesium ...

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11. Bond energy is the energy a. required to break a chemical bond. c. required to form a chemical bond. b. released when a chemical bond breaks. d. absorbed when a chemical bond forms. 12. In a molecule of fluorine, the two shared electrons give each fluorine atom \_\_\_\_ electron(s) in the outer energy level. a. 1 c. 8 b. 2 d. 32 13. electrons in

### Chapter 12 Review 2

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