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**Chapter 22 Review:  
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Nuclear Reactions,  
section 22.2

Radioactive Decay or  
Emissions Reactions:

Unstable atoms seek to  
change their number of  
protons or neutrons.

They can do this by  
high energy nuclear  
reactions.

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Nuclear Chemistry

## **Chapter 22 Review**

### **Nuclear Chemistry**

NUCLEAR CHEMISTRY

701 SECTION 22-1 O

BJECTIVES Explain

what a nuclide is, and describe the different ways nuclides can be represented. Define and relate the terms mass defect and nuclear binding energy. Explain the relationship between nucleon number and stability of nuclei.

Explain why nuclear

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Nuclear Chemistry

reactions occur, and  
know how to balance a  
nuclear ...

## **CHAPTER 22 Nuclear Chemistry**

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Nuclear Chemistry

Mixed chapter 22

review nuclear

chemistry CHAPTER 22

Nuclear Chemistry

energy levels

According to the

nuclear shell

model, nucleons exist in

different energy levels,

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Nuclear Chemistry

Section 2 Answers

Modern  
or shells, in the nucleus  
The numbers of  
nucleons that  
represent completed  
nuclear energy  
levels—2, 8, 20, 28, 50,  
82, and

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**Study ...**

fusion, controlled

nuclear reactions,

radiation. Section 22.1

Define nucleon,

isotope, nuclide, and

nuclear reaction.

Summarize the

differences between

nuclear reactions and

chemical reactions.

Section 22.2 Define

radioactivity and

radionuclide. Write

balanced equations for

nuclear reactions,

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**NUCLEAR**

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## Chapter 22 Review

### CHAPTER 22 Nuclear

### Chemistry energy

levels According to the nuclear shell

model, nucleons exist in different energy levels, or shells, in the nucleus

The numbers of nucleons that represent completed nuclear energy levels—2, 8, 20, 28, 50, 82, and 126—are called magic numbers

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Nuclear Chemistry  
**Chemistry Section 4**

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5-21 Answer Key**

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Chemistry GCC CHM

152 Nuclear chemistry

involves changes in the

nucleus (protons and

neutrons) of

radioactive atoms.

Applications of nuclear

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Nuclear Chemistry

Section 22-1: Answers

Modern

chemistry: medical

diagnosis and

treatment C-14 dating

nuclear power plants

create new elements

Nuclear Chemistry

Nuclei and Nuclear

Reactions

## **Two Types of**

## **Nuclear Processes**

Chapter 22: Nuclear

Chemistry Section

22-1: The Nucleus •

Atomic nuclei= protons

and neutrons (together

are nucleons) o

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Chapter 22 Review

Nuclear Chemistry

Solution 2 Answers

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Nuclide= an atom—identified by # of protons/neutrons in nucleus  
Mass Defect and Nuclear Stability •

Mass defect= difference between mass of an atom and sum of the masses of protons/neutrons/electrons  
o Caused by conversion of mass to energy when nucleus forms Nuclear ...

**Chapter 22 Notes -**

**Chapter 22 Nuclear**

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**Chemistry Section**

**22 ...**

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Chemistry Section 2**

...

CHAPTER 22 TEST

Nuclear Chemistry

Class MULTIPLE CHOICE

On the line at the left

of each statement,

write the letter of the

choice that best

completes the

statement or answers

the question. After

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Nuclear Chemistry

Section 22.2

converting units, the

nuclear mass defect is

equivalent to the a.

atomic mass b.

electrostatic force c.

energy of chemical

reaction

## **San Ramon Valley High School**

21.2: Patterns of  
Nuclear Stability.

21.2.1 Neutron-to-

Proton Ratio. strong

nuclear force – a strong

force of attraction

between a large

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Nuclear Chemistry

Section 22.1 Answers

Modern

number of protons in the small volume of the nucleus; stable nuclei with low atomic numbers up to 20 have nearly equal number of neutrons and protons

**21.S: Nuclear Chemistry (Summary) - Chemistry LibreTexts**

22. When a radioactive nuclide has a neutron to proton ratio that is too low, it can move

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Nuclear Chemistry

Section 2 Answers

toward stability in one of two ways, positron emission or electron capture.

## **Chapter 18 Nuclear Chemistry**

Chapter 22 Review

Organic Chemistry

Section 1 Answers

chapter 22 review

organic chemistry

section 1 answers

compilations from not

far off from the world

behind more, ...

CHAPTER 22. NUCLEAR

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University Section 224

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CHAPTER 22 Nuclear

Chemistry energy

levels According to the

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Section 3

nuclear shell  
model, nucleons exist in  
different energy levels,  
or shells, in the nucleus

The numbers of  
nucleons that  
represent completed  
nuclear energy  
levels—2, 8, 20, 28, 50,  
82, and 126—are  
called magic numbers

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Chemistry Section 3**

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