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## **Chapter 16 Review Acid Base**

Chapter 16 - Acid-Base Equilibria. 16.1

Acids & Bases: A Brief Review. -

Arrhenius acids and bases: -- acid: an

$H^+$  donor  $HA \rightleftharpoons H^+ + A^-(aq)$  (aq) (aq) -- base:

an  $OH^-$  donor  $MOH \rightleftharpoons M^+ + OH^-(aq)$  (aq) (aq) -

Brønsted-Lowry acids and bases: -- acid:

an  $H^+$  donor  $HA \rightleftharpoons H^+ + A^-(aq)$  (aq) (aq) --

base: an  $H^+$  acceptor  $HB \rightleftharpoons H^+ + B^-(aq)$  (aq)

(aq) .

## **Chapter 16 Acid-Base Equilibria - Directory**

16.2 Water and the pH Scale. It is crucial to understand the acid-base nature of

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water itself. Amphiprotic substance: A substance that can act as either an acid or a base. Water is a primary example of this. Autoionization Reaction: -A reaction that results in the formation of a cation or an anion.

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## **Chapter 16 - Acid-Base Equilibria**

Chapter 16: Acid and Base Review  
Supplemental Instruction Iowa State  
University Leader: Kelsey Course:  
Chemistry 178 Instructor: Verkade Date:  
10/10/2011 ~PLEASE DO NOT WRITE ON  
THIS WORKSHEET~ 1. What two  
substances are always produced by a  
neutralization reaction? a. acid and a  
base b. water and a base c. water and  
an acid d. water and a salt 2.

## **Chapter 16: Leader: Acid and Base**

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## **Review**

Chapter 16. 16.1 Acids and Bases: A Brief Review; 16.2 Bronsted-Lowry Acids and Bases; 16.3 The Autoionization of Water; 16.4 The pH Scale; 16.5 Strong Acids and Bases; 16.6 Weak Acids; 16.7 Weak Bases; 16.8 Relationship Between  $K_a$  and  $K_b$ ; 16.9 Acid-Base Properties of Salt Solutions; 16.10 Acid-Base Behavior and Chemical Structure; 16.11 ...

## **16.1 Acids and Bases: A Brief Review - Chapter 1 | Dr. Fus**

Chapter 16 Acids and Bases 1. Acids were recognized primarily from their sour taste. Bases were recognized from their bitter taste and slippery feel on skin.

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## **Chapter 16 (Acid-Base Equilibria) - Part 3**

Major topics: Arrhenius vs. Bronsted-Lowry definition of acids and bases, conjugate acid/base, acid dissociation constant ( $K_a$ ), & strong vs weak acids.

## **Chapter 16 (Acid-Base Equilibria) - Part 1**

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16.2.3 Related Strengths of Acids and Bases. the stronger the acid, the weaker the conjugate base. the stronger the base, the weaker the conjugate acid. equilibrium favors transfer of proton from stronger acid to stronger base.

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## **16.S: Acid-Base Equilibria**

### **(Summary) - Chemistry LibreTexts**

CHAPTER 14 REVIEW Acids and Bases

SECTION 1 SHORT ANSWER Answer the

following questions in the space

provided. 1. Name the following

compounds as acids: sulfuric acid a. H

2SO<sub>4</sub> sulfurous acid b. H<sub>2</sub>SO<sub>3</sub>

hydrosulfuric acid c. H<sub>2</sub>S perchloric acid

d. HClO<sub>4</sub> hydrocyanic acid e. hydrogen

cyanide 2. H

## **14 Acids and Bases**

Section 16.1 – ACIDS AND BASES: A

BRIEF REVIEW • Acids and bases were

first recognized by the properties of their

aqueous solutions. o For example, acids

turn litmus red, whereas bases turn

litmus blue. • Arrhenius recognized that

properties of acidic solutions are due to

H<sup>+</sup> (aq) ions and those of basic

solutions are due to OH<sup>-</sup>

## **chapter 16 - Section 16.1 ACIDS AND BASES A BRIEF REVIEW ...**

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### **Chapter 16 (Acid-Base Equilibria) - Part 2**

Chapter 16: Acid-Base Equilibria and  
Solubility Equilibria 7. You are asked to  
go into the lab and prepare an acetic  
acid - sodium acetate buffer solution  
with a pH of  $4.00 \pm 0.02$ . What molar  
ratio of  $\text{CH}_3\text{COOH}$  to  $\text{CH}_3\text{COONa}$   
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